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pH and pOH The pH of a solution indicates how acidic or basic that solution is. pH range of 0-7 acidic 7 neutral 7-14 basic Since $[H^+]\cdot[OH^-] = 10^{-14}$ at $25^\circ C$, if $[H^+]$ is known, the $[OH^-]$ can be calculated and vice versa.
 $pH = -\log [H^+]$ So if $[H^+] = 10^{-6} M$, $pH = 6$
 $pOH = -\log [OH^-]$ So if $[OH^-] =$

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10-8, $pOH = 8$ Together, $pH + pOH = 14$.

pH and pOH - Newbury Park High School

pH and pOH CONTINUED 1. 0.01 M HCl

$[H^+] = 0.01M$ $pH = -\log(0.01) = 2$ 2.

0.0010 M NaOH $[OH^-] = 0.0010M$ $pOH =$

$-\log(0.0010) = 3 \rightarrow pH = 11$ 3. 0.050 M

NaOH

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pH and pOH CONTINUED - Newbury Park High School

pH/pOH. pH and pOH are a measure of how acidic or how basic a solution is and for most solutions, is measured on a scale of 0 to 14. The lower the number, the more acidic the solution is, and the higher the number, the more basic the

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solution is. Neutral solutions (not acidic or basic) have a pH of 7.

pH / pOH - Chemistry Help

Chemistry If8766 Ph And Poh pH and pOH CONTINUED 1. 0.01 M HCl $[H^+] = 0.01M$ $pH = -\log(0.01) = 2$ 2. 0.0010 M NaOH $[OH^-] = 0.0010M$ $pOH = -\log(0.010) = 3 \rightarrow pH = 11$ 3. 0.050 M

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solution is, and the higher the number, the more basic the solution is.

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